

EXTRA PRACTICE: Naming & Writing Ionic Compounds

Name _____

Date _____ Pd _____

	Name of Cation	Name of Anion	Formula of Cation	Formula of Anion	Formula of Compound	Name of Compound
1	Calcium Cation	Chloride				
2	Iron (III) Cation	Phosphide				
3			Na ¹⁺	S ²⁻		
4			Al ³⁺	Br ¹⁻		
5						Lithium Nitride
6						Platinum (IV) Oxide
7	Magnesium Cation	Selenide				
8			Ca ²⁺	F ¹⁻		
9					HgS	
10					CuI ₂	

Name the following ionic compounds:

- 1) NaBr = _____
- 2) ScCl₃ = _____
- 3) V₂O₃ = _____
- 4) NaF = _____
- 5) Ca₃P₂ = _____
- 6) FeSe₂ = _____
- 7) Li₃P = _____
- 8) Zn₃P₂ = _____
- 9) Sr₃N₂ = _____
- 10) CuO = _____
- 11) AgI = _____
- 12) AlCl₃ = _____
- 13) SnS₂ = _____
- 14) PbF₄ = _____
- 15) K₂Se = _____
- 16) Pb₃N₂ = _____
- 17) Ni₃P₂ = _____
- 18) CdS = _____
- 19) CuCl₂ = _____
- 20) FeBr₂ = _____

Write the formulas for the following ionic compounds:

- 21) lithium iodide = _____
- 22) iron (II) sulfide = _____
- 23) titanium (II) selenide = _____
- 24) calcium bromide = _____
- 25) gallium chloride = _____
- 26) sodium nitride = _____
- 27) beryllium phosphide = _____
- 28) zinc oxide = _____
- 29) manganese (VII) selenide = _____
- 30) copper (II) chloride = _____
- 31) cobalt (III) sulfide = _____
- 32) cadmium oxide = _____
- 33) potassium nitride = _____
- 34) lead (IV) sulfide = _____
- 35) silver chloride = _____
- 36) vanadium (V) nitride = _____
- 37) strontium fluoride = _____
- 38) barium sulfide = _____
- 39) platinum (II) oxide = _____
- 40) magnesium phosphide = _____