

Unit 8 Study Guide: Solids, Liquids, Gases

1) What are 4 characteristics of gases according to the Kinetic Molecular Theory?

- | | |
|----|----|
| a) | c) |
| b) | d) |

2) Identify as pressure, volume, temperature, or moles.

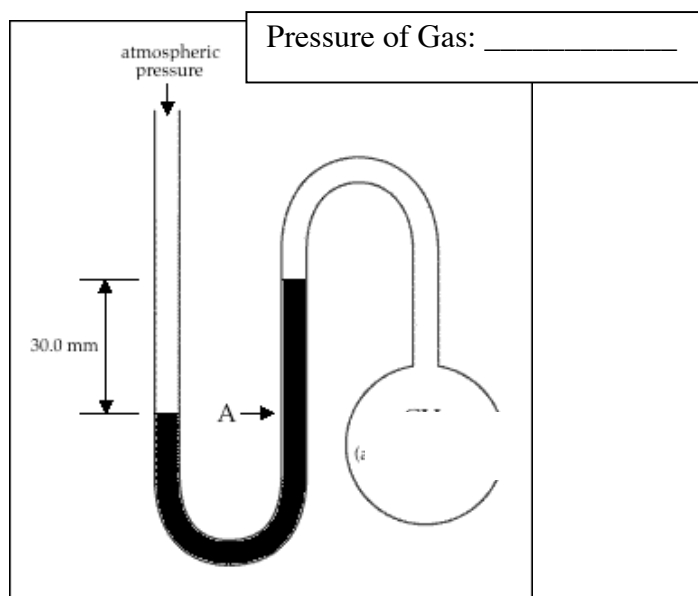
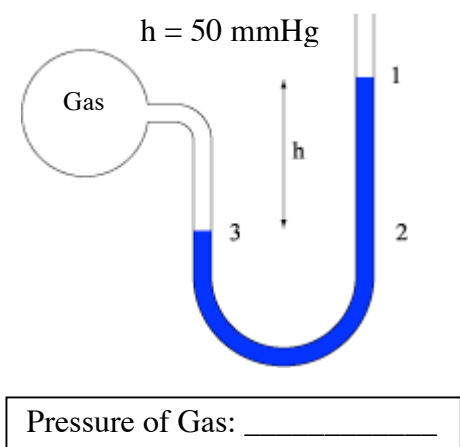
- a) The speed of the particles
- b) The size of the container
- c) The number of collisions between the particles and the walls of the container
- d) The number of particles in the container

3) Match the law with the equation:

_____ Boyle's Law	a) $PV = nRT$	d) $P_1V_1 = P_2V_2$
_____ Charles's Law	b) $\frac{P_1}{T_1} = \frac{P_2}{T_2}$	e) $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$
_____ Gay-Lussac's Law	c) $\frac{V_1}{T_1} = \frac{V_2}{T_2}$	f) $\frac{V_1}{n_1} = \frac{V_2}{n_2}$
_____ Avogadro's Law		
_____ Combined Gas Law		
_____ Ideal Gas Law		

4) **Reading a Manometer:**

Determine the pressure of the gas if the atmospheric pressure is 750 mmHg for both.



5) Convert these pressures. SHOW YOUR WORK.

- | | |
|-------------------------|----------------------------|
| a) 600 Torr = _____ atm | c) 2.5 atm = _____ kPa |
| b) 40 kPa = _____ mmHg | d) 755 mmHg = _____ Torr ☺ |

Gas Law Problems:

- 6) If the volume of a gas is expanded from 45.0 mL to 105 mL, what is the new temperature if the temperature starts at 15°C and the pressure is held constant?

Equation: _____ Solution: _____

- 7) If the pressure on a gas is increased from 750 mmHg to 900 mmHg, what is the new volume if it was initially 250 mL and the temperature is held constant?

Equation: _____ Solution: _____

- 8) What number of moles of carbon dioxide gas occupies 55.5 L at 95 kPa and 20°C?

Equation: _____ Solution: _____

- 9) If the pressure on a sample of gas decreases from 1.25 atm to 0.75 atm and the volume expands from 2.44 L to 2.90 L, what is the final temperature if the initial temperature of the gas is -40°C?

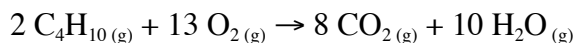
Equation: _____ Solution: _____

- 10) If the temperature on a sample of gas increases from 0°C to 37°C, what is the new pressure if the initial pressure was 1.00 atm?

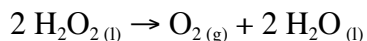
Equation: _____ Solution: _____

Gas Stoichiometry & Gas Laws:

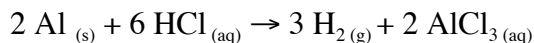
- 11) What are the values for standard temperature and pressure?
- 12) (*Easiest*) Given the reaction shown here, how many liters of CO₂ will be produced from the combustion of 0.750 L of C₄H₁₀ if this reaction occurs at constant temperature and pressure.



- 13) (*Medium*) If 25 grams of hydrogen peroxide decompose at STP according to the following reaction, how many liters of oxygen are produced?

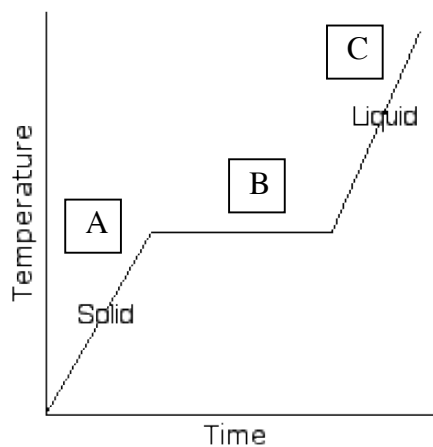


- 14) (*Hardest*) Given the reaction of 0.25 g of Al with HCl, how many liters of H₂ gas will be produced if the temperature is 40°C and the pressure is 1.15 atm? *Hint: Is this at STP?*



Lab Type Questions:

- 15) You collect a gas over water at 20°C and the vapor pressure of the water is 21.1 mmHg. What is the pressure on the gas you collected if the atmospheric pressure outside the gas collection chamber is 755 mmHg. Express your answer in atmospheres.
- 16) You react a strip of magnesium with HCl in a gas collection tube over water. The temperature of the water and gas are 22°C. The vapor pressure of water at that temperature is 19.8 mmHg and the atmospheric pressure is 730 mmHg. The volume of gas you collect is 92.5 mL. What mass of magnesium must you use in the reaction?
Hint: very similar to #15, but you have to remove water vapor first!



- 17) To the left is shown a heating curve for a substance. Describe what is happening in the three parts of the graph as heat is added. Be sure to include: temperature, kinetic energy, and potential energy.

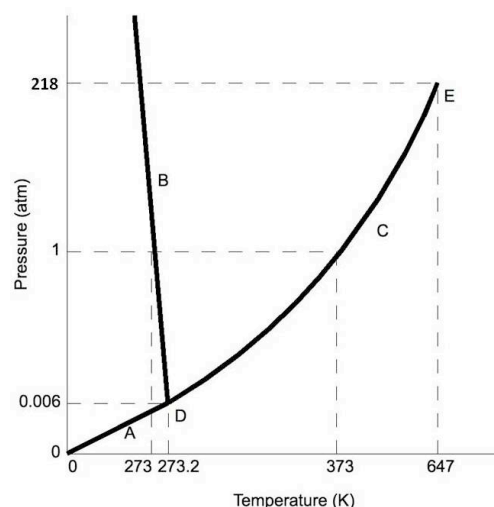
A:

B:

C:

18) Answer the following questions using the phase diagram shown here.

- Could this be a phase diagram for water?
- What is the approximate value for the critical temperature?
- Label the three phases on the diagram.
- Along which line does freezing occur?
- What is the temperature and pressure at the triple point?
- Sketch a heating curve for this substance going from 0 K to 647 K at 1 atm.



Intermolecular (Interparticle) Forces (IMF):

19) Identify type of substance as ionic, polar covalent, non-polar covalent, network covalent, or metallic. Identify the interparticle force (IMF) that holds the substance together: hydrogen bond, ion-dipole, London dispersion force, metallic, ionic, or dipole-dipole.

Substance	Type of Substance	Interparticle Force
a) SiO_2 (quartz)		
b) NBr_3		
c) Fe		
d) ZnCl_2		
e) SO_2		
f) Br_2		
g) HBr		

20) Of the substances in the table above:

- Which should have the highest melting point?
- Which should have the lowest boiling point?
- Which should be the most ductile?
- Which has a bond that involves sharing electrons unequally?
- Which is made of cations that are bonded together by a sea of electrons?
- Which is bonded together by a transfer of electrons?
- Which would be considered an electrolyte?