

# Double Replacement Reactions LAB

Name: \_\_\_\_\_

## PROCEDURES:

- 1. For each reaction, combine **TWO** drops of **each** reactant onto the wax paper (**DO NOT** allow tips of droppers to touch solutions). **CIRCLE** reaction/no reaction based on your observations (some reactions should be observed **VERY** carefully).
  - If a reaction occurs, record whether a **PRECIPITATE** is formed on the line, and if so, record the precipitate **COLOR**.
- 2. **AFTER** running all reactions, predict the products and its state of matter (*aqueous or solid*) and then balance the equation.
- 3. In each balanced chemical equation, **CIRCLE** which product is the precipitate (*solid*), if there is one produced.
- 4. For **each** reaction, write its **COMPLETE (TOTAL) IONIC AND NET IONIC EQUATION**.
- 5. If no reaction occurs, write its complete (total) ionic equation, but in the **net ionic equation**, write **"NO REACTION"**.

1) **CIRCLE ONE:** Reaction / No Reaction      **OBSERVATIONS:** \_\_\_\_\_



\*\* Complete (total) Ionic Equation:

\*\* Net Ionic Equation:

2) **CIRCLE ONE:** Reaction / No Reaction      **OBSERVATIONS:** \_\_\_\_\_



\*\* Complete (total) Ionic Equation:

\*\* Net Ionic Equation:

3) **CIRCLE ONE:** Reaction / No Reaction      **OBSERVATIONS:** \_\_\_\_\_



\*\* Complete (total) Ionic Equation:

\*\* Net Ionic Equation:

4) CIRCLE ONE: Reaction / No Reaction      OBSERVATIONS: \_\_\_\_\_

Reactants: \_\_\_\_\_  $\text{Pb}(\text{NO}_3)_2$  ( ) + \_\_\_\_\_  $\text{NaOH}$  ( )  $\rightarrow$

\*\* Complete (total) Ionic Equation:

\*\* Net Ionic Equation:

5) CIRCLE ONE: Reaction / No Reaction      OBSERVATIONS: \_\_\_\_\_

Reactants: \_\_\_\_\_  $\text{CoCl}_2$  ( ) + \_\_\_\_\_  $\text{KI}$  ( )  $\rightarrow$

\*\* Complete (total) Ionic Equation:

\*\* Net Ionic Equation:

6) CIRCLE ONE: Reaction / No Reaction      OBSERVATIONS: \_\_\_\_\_

Reactants: \_\_\_\_\_  $\text{CuCl}_2$  ( ) + \_\_\_\_\_  $\text{NaOH}$  ( )  $\rightarrow$

\*\*\*Complete (total) Ionic Equation:

\*\*\*Net Ionic Equation:

7) CIRCLE ONE: Reaction / No Reaction      OBSERVATIONS: \_\_\_\_\_

Reactants: \_\_\_\_\_  $\text{Pb}(\text{NO}_3)_2$  ( ) + \_\_\_\_\_  $\text{KI}$  ( )  $\rightarrow$

\*\*\*Complete (total) Ionic Equation:

\*\*\*Net Ionic Equation: