

Double Replacement Reactions

Name: _____

Block: _____

Date: _____

In this lab we will look at 7 possible double replacement reactions and investigate why some reactions occur and others do not.

Procedure:

- Write down the formulas of the reactants and then use the back of your periodic table to figure out what ions (with their charges) make up those compounds.
- Next, try the reaction by combining a drop or two of each reactant at your lab station. If a reaction occurs, make a note of what happens.
- Use your reference tables to predict the products of each reaction and assess the solubility of each product.
- Write balanced equations for each reaction that occurs.

1) Reactants: cobalt (II) chloride and sodium phosphate

Ions in solution:

Observations: _____ **Reaction** **No Reaction**

What are the names of the two products formed from the ions listed above? Circle any product that is insoluble in water.

_____ and _____

Write a balanced equation for the reaction:

Write the net ionic equation for the reaction:

2) Reactants: copper (II) chloride and sodium phosphate

Ions in solution:

Observations: _____ **Reaction** **No Reaction**

What are the names of the two products formed from the ions listed above? Circle any product that is insoluble in water.

_____ and _____

Write a balanced equation for the reaction:

Write the net ionic equation for the reaction:

3) Reactants: silver nitrate and sodium hydroxide

Ions in solution:

Observations: _____ **Reaction** **No Reaction**

What are the names of the two products formed from the ions listed above? Circle any product that is insoluble in water.

_____ and _____

Write a balanced equation for the reaction:

Write the net ionic equation for the reaction:

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4) Reactants: lead (II) nitrate and sodium hydroxide ions in solution:

Observations: _____ Reaction No Reaction

What are the names of the two products formed from the ions listed above? Circle any product that is insoluble in water.

_____ and _____

Write a balanced equation for the reaction:

Write the net ionic equation for the reaction:

5) Reactants: cobalt (II) chloride and potassium iodide ions in solution:

Observations: _____ Reaction No Reaction

What are the names of the two products formed from the ions listed above? Circle any product that is insoluble in water.

_____ and _____

Write a balanced equation for the reaction:

Write the net ionic equation for the reaction:

6) Reactants: copper (II) chloride and sodium hydroxide ions in solution:

Observations: _____ Reaction No Reaction

What are the names of the two products formed from the ions listed above? Circle any product that is insoluble in water.

_____ and _____

Write a balanced equation for the reaction:

Write the net ionic equation for the reaction:

7) Reactants: lead (II) nitrate and potassium iodide ions in solution:

Observations: _____ Reaction No Reaction

What are the names of the two products formed from the ions listed above? Circle any product that is insoluble in water.

_____ and _____

Write a balanced equation for the reaction:

Write the net ionic equation for the reaction: