

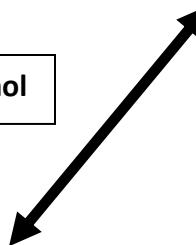
Volume  
of gas



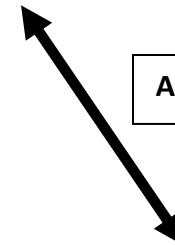
1 mole = 22.4 Liters (L) of gas @ STP

MOLE

Molar Mass → g/mol



Avogadro's # →  $6.02 \times 10^{23}$  Particles



Mass

Particles

- 1 mole =  $6.02 \times 10^{23}$  Particles (atoms, molecules, formula units)
- 1 mole = Molar Mass (g/mol)
- 1 mole = 22.4 L of gas @ STP (Standard Temperature Pressure)

# MOLE-ville

