

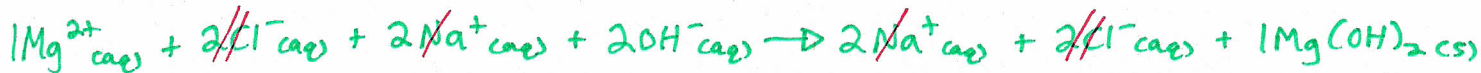
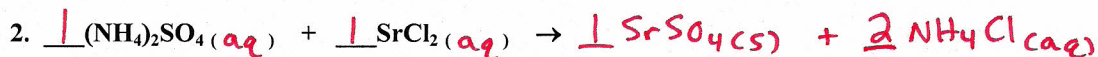
EXTRA PRACTICE: Net Ionic Equations Practice #3

Name: _____

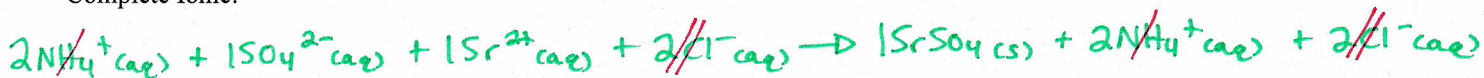
Predict the products of the following reactions and then balance. Determine if each compound is aqueous (aq) or solid (s) using the solubility rules provided for you. Write out the full complete ionic equation, spectator ions, and net ionic equation.



Complete Ionic:

Spectator Ions: $2\text{Na}^{+}(\text{aq})$; $2\text{Cl}^{-}(\text{aq})$ 

Complete Ionic:

Spectator Ions: $2\text{NH}_4^{+}(\text{aq})$; $2\text{Cl}^{-}(\text{aq})$ 

Complete Ionic:

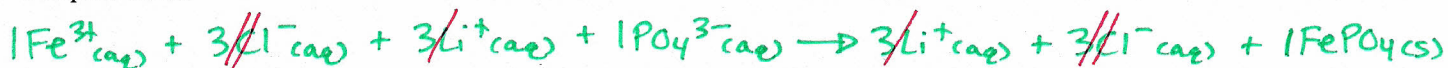


Spectator Ions: ALL

Net Ionic: (NR)



Complete Ionic:

Spectator Ions: $3\text{Li}^{+}(\text{aq})$; $3\text{Cl}^{-}(\text{aq})$ 

Complete Ionic:

Spectator Ions: $2\text{NH}_4^{+}(\text{aq})$; $2\text{NO}_3^{-}(\text{aq})$ 