

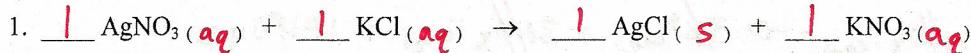
# ANSWER KEY

## EXTRA PRACTICE: Net Ionic Equations Practice #2

Name: \_\_\_\_\_

Complete the following net ion equations. Be sure to show **ALL** of your work for full credit. Balance **ALL** equations first.

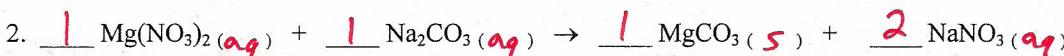
If no reaction occurs, you must still balance the equation, but write **NR** under net ionic equation. Be sure to also carry through your charges and coefficients into the final net ionic equation. **Hint: Determine if each compound is aqueous (aq) or not first.**



Complete Ionic:



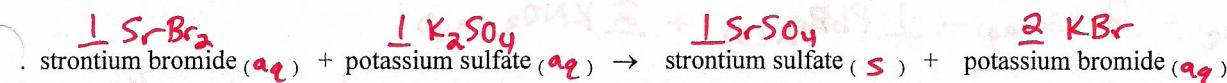
Spectator Ions:  $1\text{K}^+(\text{aq})$ ;  $1\text{NO}_3^-(\text{aq})$



Complete Ionic:



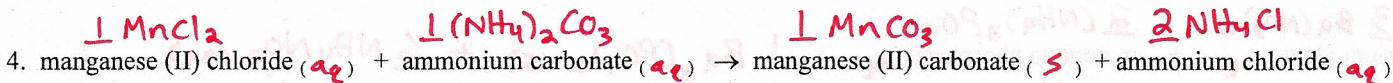
Spectator Ions:  $2\text{Na}^+(\text{aq})$ ;  $2\text{NO}_3^-(\text{aq})$



Complete Ionic:



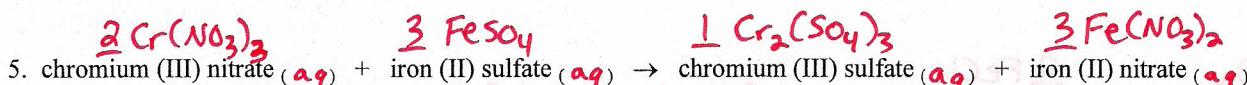
Spectator Ions:  $2\text{K}^+(\text{aq})$ ;  $2\text{Br}^-(\text{aq})$



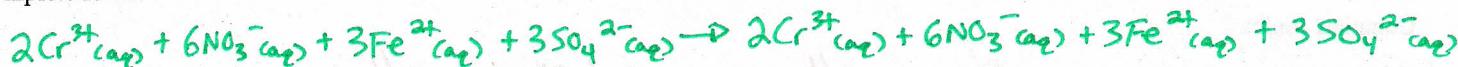
Complete Ionic:



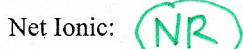
Spectator Ions:  $2\text{NH}_4^+(\text{aq})$ ;  $2\text{Cl}^-(\text{aq})$



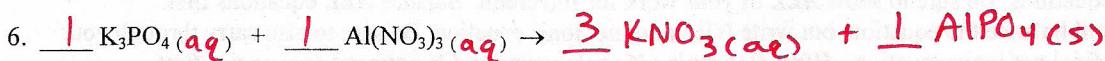
Complete Ionic:



Spectator Ions: **ALL**



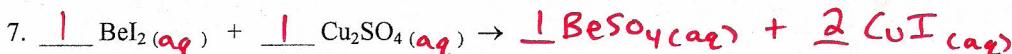
Predict the products of the following reactions, and show the complete ionic, spectator ions, and net ionic forms of the equation:



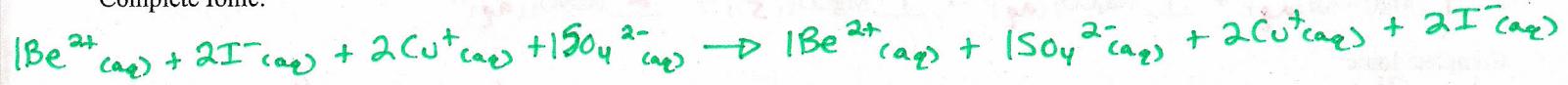
Complete Ionic:



Spectator Ions:  $3K^+(aq)$  ;  $3NO_3^-(aq)$



Complete Ionic:



Spectator Ions: ALL

Net Ionic: (NR)



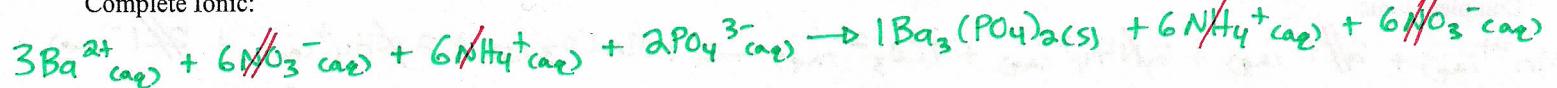
Complete Ionic:



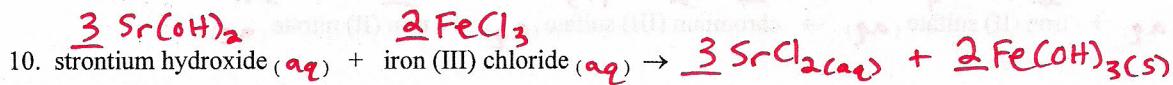
Spectator Ions:  $2K^+(aq)$  ;  $2NO_3^-(aq)$



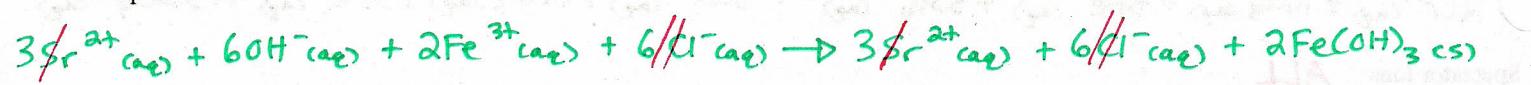
Complete Ionic:



Spectator Ions:  $6NH_4^+(aq)$  ;  $6NO_3^-(aq)$



Complete Ionic:



Spectator Ions:  $3Sr^{2+}(aq)$  ;  $6Cl^-(aq)$

