ANSWER KEY

EXTRA PRACTICE: Naming & Writing Ionic Compounds

Name _____

Date _____ Pd _____

	Name of Cation	Name of Anion	Formula of Cation	Formula of Anion	Formula of Compound	Name of Compound
1	Calcium Cation	Chloride	Ca^{2+}	Cl^{-}	$CaCl_2$	Calcium Chloride
2	Iron (III) Cation	Phosphide	Fe^{3+}	P^{3-}	FeP	Iron (III) Phosphide
3	Sodium Cation	Sulfide	Na^{1+}	S^{2-}	Na_2S	Sodium Sulfide
4	Aluminum Cation	Bromide	Al^{3+}	Br^{1-}	$AlBr_3$	Aluminum Bromide
5	Lithium Cation	Nitride	Li^{+}	N^{3-}	Li_3N	Lithium Nitride
6	Platinum (IV) Cation	Oxide	Pt^{4+}	O^{2-}	PtO_2	Platinum (IV) Oxide
7	Magnesium Cation	Selenide	Mg^{2+}	Se^{2-}	$MgSe$	Magnesium Selenide
8	Calcium Cation	Fluoride	Ca^{2+}	F^{1-}	CaF_2	Calcium Fluoride
9	Mercury (II) Cation	Sulfide	Hg^{2+}	S^{2-}	HgS	Mercury (II) Sulfide
10	Copper (II) Cation	Iodide	Cu^{2+}	I^{-}	CuI_2	Copper (II) Iodide

Name the following ionic compounds:

1) $NaBr$ = Sodium Bromide

2) $ScCl_3$ = Scandium (III) Chloride

3) V_2O_3 = Vanadium (III) Oxide

4) NaF = Sodium Fluoride

5) Ca_3P_2 = Calcium Phosphide

6) $FeSe_2$ = Iron (IV) Selenide

7) Li_3P = Lithium Phosphide

8) Zn_3P_2 = Zinc Phosphide

9) Sr_3N_2 = Strontium Nitride

10) CuO = Copper (II) Oxide

11) AgI = Silver Iodide

12) $AlCl_3$ = Aluminum Chloride

13) SnS_2 = Tin (IV) Sulfide

14) PbF_4 = Lead (IV) Fluoride

15) K_2Se = Potassium Selenide

16) Pb_3N_2 = Lead (II) Nitride

17) Ni_3P_2 = Nickel (II) Phosphide

18) CdS = Cadmium Sulfide

19) $CuCl_2$ = Copper (II) Chloride

20) $FeBr_2$ = Iron (II) Bromide

Write the formulas for the following ionic compounds:

21) Li^{+1} I^{-1} lithium iodide = LiI

22) Fe^{2+} S^{2-} iron (II) sulfide = FeS

23) Ti^{2+} Se^{2-} titanium (II) selenide = $TiSe$

24) Ca^{2+} Br^{-} calcium bromide = $CaBr_2$

25) Ga^{3+} Cl^{-} gallium chloride = $GaCl_3$

26) Na^{+} N^{3-} sodium nitride = Na_3N

27) Be^{2+} P^{3-} beryllium phosphide = Be_3P_2

28) Zn^{2+} O^{2-} zinc oxide = ZnO

29) Mn^{7+} Se^{2-} manganese (VII) selenide = Mn_2Se_7

30) Cu^{2+} Cl^{-} copper (II) chloride = $CuCl_2$

31) Co^{3+} S^{2-} cobalt (III) sulfide = Co_2S_3

32) Cd^{2+} O^{2-} cadmium oxide = CdO

33) K^{+} N^{3-} potassium nitride = K_3N

34) Pb^{4+} S^{2-} lead (IV) sulfide = PbS_2

35) Ag^{+} Cl^{-} silver chloride = $AgCl$

36) V^{5+} N^{3-} vanadium (V) nitride = V_3N_5

37) Sr^{2+} F^{-} strontium fluoride = SrF_2

38) Ba^{2+} S^{2-} barium sulfide = BaS

39) Pt^{2+} O^{2-} platinum (II) oxide = PtO

40) Mg^{2+} P^{3-} magnesium phosphide = Mg_3P_2