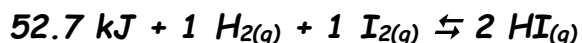


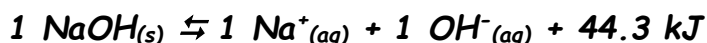
**EXTRA PRACTICE: Le Chatelier's Principle Practice #3** Name: \_\_\_\_\_

Part I: Complete the following table. Write (*left, right, or none*) for equilibrium shift, and (*decreases, increases, or remains same*), for concentrations of reactants and products and for equilibrium constant (K):



Stress Type	Equilibrium Shift	[H <sub>2</sub> ]	[I <sub>2</sub> ]	[HI]	Equilibrium Constant (K)
Add H <sub>2</sub>		-----			
Add I <sub>2</sub>			-----		
Add HI				-----	
Remove H <sub>2</sub>		-----			
Remove I <sub>2</sub>			-----		
Remove HI				-----	
↑ Temperature					
↓ Temperature					
↑ Pressure					
↓ Pressure					

Part II: Complete the following table. Write (*left, right, or none*) for equilibrium shift, and (*decreases, increases, or remains same*), for concentrations of reactants and products and for equilibrium constant (K):



(Remember that pure solids and liquids do not affect equilibrium values)

Stress Type	Equilibrium Shift	Amount of NaOH <sub>(s)</sub>	[Na <sup>+</sup> ]	[OH <sup>-</sup> ]	Equilibrium Constant (K)
Add NaOH <sub>(s)</sub>		-----			
Add NaCl (Adds Na <sup>+</sup> )			-----		
Add KOH (Adds OH <sup>-</sup> )				-----	
Add H <sup>+</sup> (Removes OH <sup>-</sup> )				-----	
↑ Temperature					
↓ Temperature					
↑ Pressure					
↓ Pressure					