

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Pd: \_\_\_\_\_

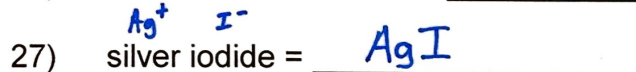
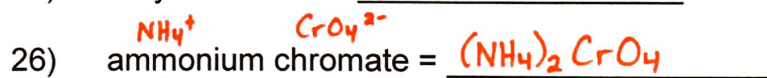
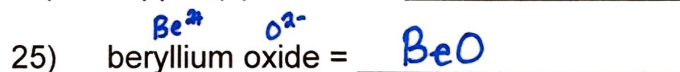
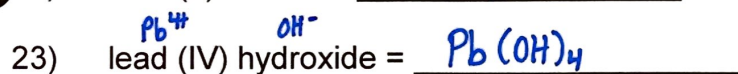
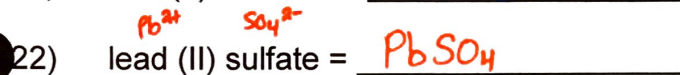
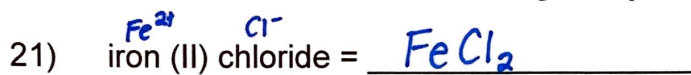
**EXTRA PRACTICE: Naming & Writing Ionic Compounds with Polyatomics****Name the following ionic compounds:**

- 1)  $\text{NH}_4\text{Cl} =$  Ammonium Chloride
- 2)  $\text{Fe}(\text{NO}_3)_3 =$  Iron (III) Nitrate
- 3)  $\text{TiBr}_3 =$  Titanium (III) Bromide
- 4)  $\text{Cu}_3\text{P} =$  Copper (I) Phosphide
- 5)  $\text{SnSe}_2 =$  Tin (IV) Selenide
- 6)  $\text{GaN} =$  Gallium Nitride
- 7)  $\text{Pb}(\text{SO}_4)_2 =$  Lead (IV) Sulfate
- 8)  $\text{Be}(\text{HCO}_3)_2 =$  Beryllium Bicarbonate
- 9)  $\text{Mn}_2(\text{SO}_3)_3 =$  Manganese (III) Sulfite
- 10)  $\text{Al}(\text{CN})_3 =$  Aluminum Cyanide

**Write the formulas for the following compounds:**

- 11) chromium (VI) phosphate =  $\overset{\text{Cr}^{6+}}{\text{Cr}}(\overset{\text{PO}_4^{3-}}{\text{PO}_4})_2$
- 12) vanadium (IV) carbonate =  $\overset{\text{V}^{4+}}{\text{V}}(\overset{\text{CO}_3^{2-}}{\text{CO}_3})_2$
- 13) tin (II) nitrite =  $\overset{\text{Sn}^{2+}}{\text{Sn}}(\overset{\text{NO}_2^-}{\text{NO}_2})_2$
- 14) cobalt (III) oxide =  $\overset{\text{Co}^{3+}}{\text{Co}}_2\overset{\text{O}^{2-}}{\text{O}}_3$
- 15) titanium (II) acetate =  $\overset{\text{Ti}^{2+}}{\text{Ti}}(\overset{\text{C}_2\text{H}_3\text{O}_2^-}{\text{C}_2\text{H}_3\text{O}_2})_2$
- 16) vanadium (V) sulfide =  $\overset{\text{V}^{5+}}{\text{V}}_2\overset{\text{S}^{2-}}{\text{S}}_5$
- 17) chromium (III) hydroxide =  $\overset{\text{Cr}^{3+}}{\text{Cr}}(\overset{\text{OH}^-}{\text{OH}})_3$
- 18) lithium iodide =  $\overset{\text{Li}^+}{\text{Li}}\overset{\text{I}^-}{\text{I}}$
- 19) lead (II) nitride =  $\overset{\text{Pb}^{2+}}{\text{Pb}}_3\overset{\text{N}^{3-}}{\text{N}}_2$
- 20) silver bromide =  $\overset{\text{Ag}^+}{\text{Ag}}\overset{\text{Br}^-}{\text{Br}}$

• Write the formulas for the following compounds:



Name the following ionic compounds:

