

EXTRA PRACTICE: Naming & Writing Ionic Compounds with Polyatomics***Name the following ionic compounds:***

- 1) NH_4Cl = _____
- 2) $\text{Fe}(\text{NO}_3)_3$ = _____
- 3) TiBr_3 = _____
- 4) Cu_3P = _____
- 5) SnSe_2 = _____
- 6) GaN = _____
- 7) $\text{Pb}(\text{SO}_4)_2$ = _____
- 8) $\text{Be}(\text{HCO}_3)_2$ = _____
- 9) $\text{Mn}_2(\text{SO}_3)_3$ = _____
- 10) $\text{Al}(\text{CN})_3$ = _____

Write the formulas for the following compounds:

- 11) chromium (VI) phosphate = _____
- 12) vanadium (IV) carbonate = _____
- 13) tin (II) nitrite = _____
- 14) cobalt (III) oxide = _____
- 15) titanium (II) acetate = _____
- 16) vanadium (V) sulfide = _____
- 17) chromium (III) hydroxide = _____
- 18) lithium iodide = _____
- 19) lead (II) nitride = _____
- 20) silver bromide = _____

Write the formulas for the following compounds:

- 21) iron (II) chloride = _____
- 22) lead (II) sulfate = _____
- 23) lead (IV) hydroxide = _____
- 24) copper (II) acetate = _____
- 25) beryllium oxide = _____
- 26) ammonium chromate = _____
- 27) silver iodide = _____
- 28) potassium nitrite _____

Name the following ionic compounds:

- 29) KI = _____
- 30) MgSO_3 = _____
- 31) SnBr_4 = _____
- 32) Mn_3P_2 = _____
- 33) NaF = _____
- 34) $\text{Sr}(\text{MnO}_4)_2$ = _____
- 35) $\text{Cr}_3(\text{PO}_4)_2$ = _____
- 36) $\text{Al}(\text{ClO}_2)_3$ = _____