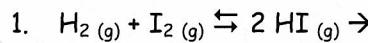


ANSWER KEY

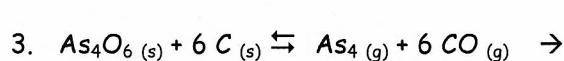
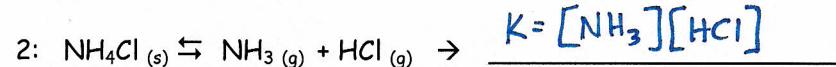
EXTRA PRACTICE: Chemical Equilibrium Practice #2

Name: _____

Write the equilibrium expression (K) for the following reactions:

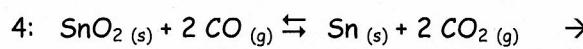


$$K = \frac{[\text{HI}]^2}{[\text{H}_2][\text{I}_2]}$$

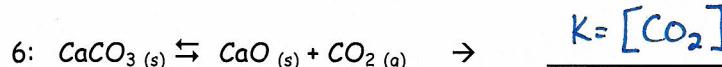


$$K = [\text{As}_4][\text{CO}]^6$$

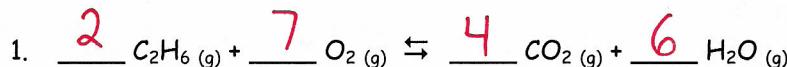
$$K = \frac{[\text{CO}_2]^2}{[\text{CO}]^2}$$



$$K = \frac{[\text{H}_2\text{O}]^3}{[\text{H}_2]^3}$$

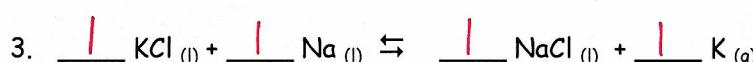
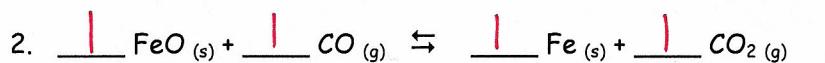


Balance each equation and then write the equilibrium expression (K) for each reaction:

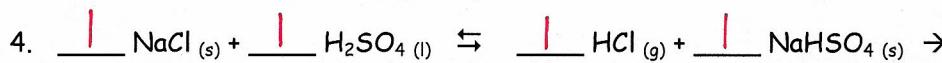


$$K = \frac{[\text{CO}_2]^4[\text{H}_2\text{O}]^6}{[\text{C}_2\text{H}_6]^2[\text{O}_2]^7}$$

$$K = \frac{[\text{CO}_2]}{[\text{CO}]}$$

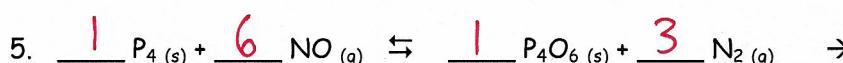


$$K = [\text{K}]$$



$$K = [\text{HCl}]$$

$$K = \frac{[\text{Na}]^3}{[\text{NO}]^6}$$



$$K = \frac{[\text{N}_2]}{[\text{NO}]^2[\text{H}_2]^2}$$

