

**EXTRA PRACTICE: Bohr Models/Electron-Dots/Valence Electrons** Name: \_\_\_\_\_

Draw the Bohr Model for each of the following elements using the **SMALLER** font element symbol on left. Then draw the Electron-Dot Structure for each element using the **LARGER** font element symbol on right.

1.

Li

**Li**

2.

Be

**Be**

3.

B

**B**

4.

C

**C**

5.

N

**N**

6.

O

**O**

7.

F

**F**

8.

Ne

**Ne**

9.

Ar

**Ar**

10.

Se

**Se**

Write the electron configuration notation (long hand method) for each of the following elements. Circle the identity of the valence electrons and determine the number of valence electrons.

1. Francium = \_\_\_\_\_

Valence Electrons = \_\_\_\_\_

2. Tin = \_\_\_\_\_

Valence Electrons = \_\_\_\_\_

3. Strontium = \_\_\_\_\_

Valence Electrons = \_\_\_\_\_

4. Phosphorus = \_\_\_\_\_

Valence Electrons = \_\_\_\_\_

5. Arsenic = \_\_\_\_\_

Valence Electrons = \_\_\_\_\_

6. Potassium = \_\_\_\_\_

Valence Electrons = \_\_\_\_\_

7. Magnesium = \_\_\_\_\_

Valence Electrons = \_\_\_\_\_

8. Iodine = \_\_\_\_\_

Valence Electrons = \_\_\_\_\_

9. Gallium = \_\_\_\_\_

Valence Electrons = \_\_\_\_\_

10. Bromine = \_\_\_\_\_

Valence Electrons = \_\_\_\_\_