Net	Ionic	Ea	uations	Pract	ice	#1
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Net Ionic:

Complete the following net ion equations. Be sure to show <i>ALL</i> of your work for full credit. Balance <i>ALL</i> equations first. If no reaction occurs, you must still balance the equation, but write NR under net ionic equation. Be sure to also carry through you charges and coefficients into the final net ionic equation. <i>Hint: Determine if each compound is aqueous (aq) or not first.</i>
1. $AgNO_3() + KCl() \rightarrow AgCl() + KNO_3()$
Complete (Total) Ionic:
Spectator Ions:
Net Ionic:
2. $Mg(NO_3)_2$ () + Na_2CO_3 () \rightarrow $MgCO_3$ () + $NaNO_3$ ()
Complete (Total) Ionic:
Spectator Ions:
Net Ionic:
3MnCl ₂ () +(NH ₄) ₂ CO ₃ () \rightarrow MnCO ₃ () +NH ₄ Cl ₁ ()
Complete (Total) Ionic:
Spectator Ions:

Name: _____

