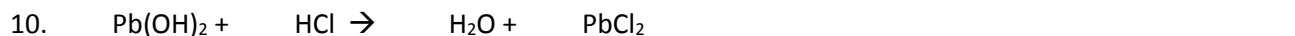


Chapter 10 - Reactions Cumulative Practice

Name: _____

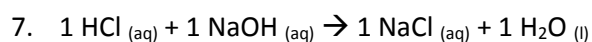
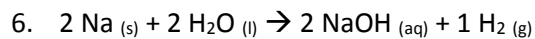
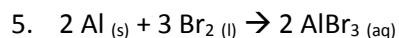
Balance each chemical equation and then identify the type of chemical reaction that is displayed.



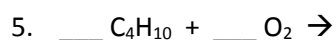
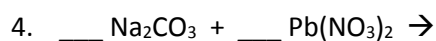
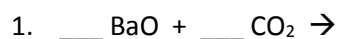
From the word equation, write a balanced chemical equation and include all states of matter (*s, l, g, aq*). From the chemical equation, write a word equation including the states of matter (*s, l, g, aq*).

- Hydrogen gas and oxygen gas chemically mix in solution to produce liquid water.
- Zinc metal reacts with liquid hydrochloric acid to produce aqueous zinc chloride and hydrogen gas.
- Aqueous Iron (III) Chloride reacts in solution with aqueous sodium hydroxide to produce the precipitate, Iron (III) hydroxide and aqueous sodium chloride.

4. Methane gas (CH₄) burns in the presence of oxygen gas to produce a combustible reaction.



Given the reactants, predict products of the following chemical reactions, and then balance the chemical equation.



Using the activity series of metals chart, determine which of the following combinations of reactants will react. Circle the pair that will react.

a) Ag and NaNO₃ / Al and NiCl₂

e) Ni and Li₃N / K and NaCl

b) Sr and FeBr₃ / Co and CaCO₃

f) Bi and CoCl₃ / Ni and CuSO₄

c) Cr and H₂SO₄ / Fe and Mg₃(PO₄)₂

g) Cd and AgNO₃ / Ag and Mg(ClO₃)₂

d) Cu and Pb(ClO₃)₂ / Mn and SnF₄

h) Sn and CaCl₂ / Ba₃P₂ and Rb