

## Unit 7 – Ch 12 – Percent Yield and Percent Error

### THEORETICAL YIELD:

➤ DEFINITION –

### ACTUAL YIELD:

➤ DEFINITION –

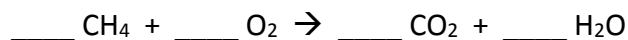
### % YIELD and % ERROR:

➤ FORMULA:            *% Yield* =

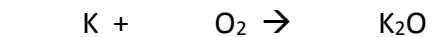
➤ FORMULA:            *% Error* =

### PRACTICE EXAMPLES:

Ex #1) Determine the *percent yield* of a reaction if 10.0 grams of methane (CH<sub>4</sub>) gas burns in excess of oxygen to experimentally produce 19.8 grams of water.



Ex #2) Determine the *percent yield* if 6.92 grams of potassium reacts with oxygen gas where 7.36 grams of potassium oxide is recovered.



\*\* Ex #3) Determine the *actual yield* if 82.4 grams of rubidium reacts with 11.6 grams of oxygen gas where 44.1% of rubidium oxide is recovered.

