

Unit 6 – Ch 11 – Molecular Formulas (M.F.)

REVIEW OF EMPIRICAL FORMULAS (E.F.):

- **EMPIRICAL FORMULAS (E.F.)** represent the _____ **AND** _____ ratio of **atoms** in the compound.
 - May be the _____ as the _____.

CALCULATING MOLECULAR FORMULAS (M.F.):

➤ DEFINITION →

- Ex: _____ → _____ atoms ; _____ atoms ; _____ atoms
- Step #1: Must have the **CORRECT** _____.
 - _____ or _____.
- Step #2: Calculate the _____ of the _____.
- Step #3: Calculate the _____.
 - **FORMULA:**
- Step #4: Multiply the _____ by the _____ calculated in step #3.

PRACTICE EXAMPLES:

1. Determine molecular formula of a compound whose molecular mass is 180 g/mol and E.F. is CH₂O.

2. A compound is 50.52 % C, 5.26 % H, and 44.22 % N. If molecular mass is 380.2 g/mol, what is the M.F.?